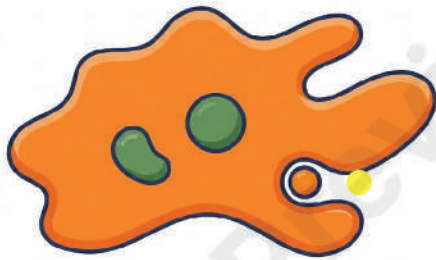
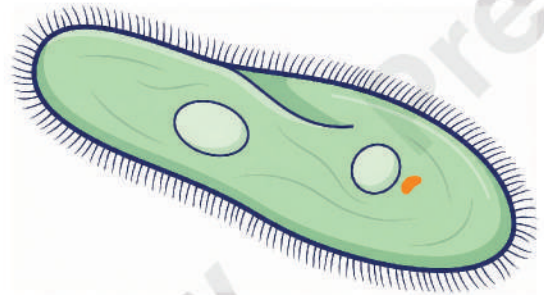


E. Amoeba and Paramecium -



Amoeba



Paramecium

Movement

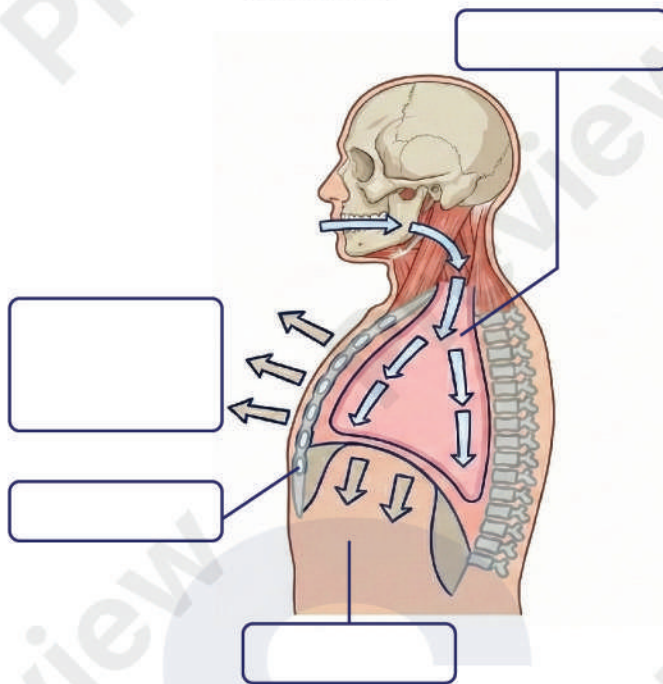
Formation of

Digestion & Absorption

Egestion

04. Breathing Mechanism-

Inhalation

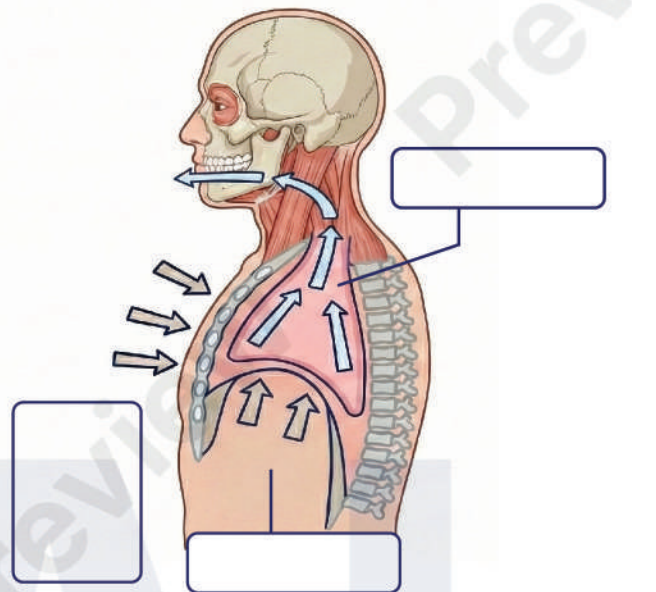


Ribs -

Diaphragm -

Chest cavity -

Exhalation



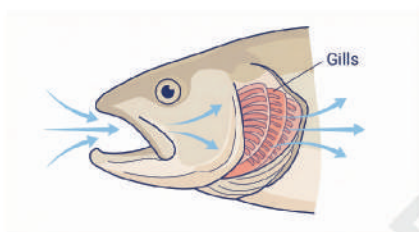
Ribs -

Diaphragm -

Chest cavity -

05. Rate of Breathing in-

Aquatic Animals

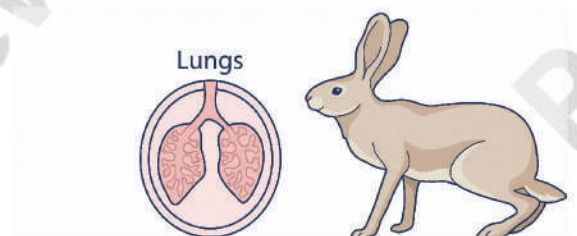


Lives in -

Amount of O_2 -

Breathing rate -

Terrestrial Animals



Lives in -

Amount of O_2 -

Breathing rate -

G. Reaction Based Conceptual Questions

Q1. A brown substance X on heating in air forms a substance Y. When hydrogen gas is passed over heated Y, it again changes back into X.

- Name X & Y
- Write chemical equations

X =	y =

Q2. A metal salt MX when exposed to light splits up to form metal M and gas X. Metal M is used in ornaments and X is used in bleaching powder. The salt MX is used in black & white photography.

- Identify MX, M & X
- Write chemical equation

MX =	M =	X =

Q3. Complete the following questions -

$\text{BaCl}_2 + \text{X} \rightarrow \text{BaSO}_4$	
$\text{Cu} + \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Y}$	
$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$	

Q4. A metal M reacts with an acid according to equation: $\text{M} + \text{H}^+ \rightarrow \text{M}^{2+} + \text{H}_2$ -

- Identify MX, M & X

M =

Q5. You heat a white powder in a test tube. A gas released turns lime water milky. A solid residue remains behind.

- Name the type of reaction
- Name of the Chemical Compound
- Name of the Gas Formed

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B. Reactions of Acids — Concepts & Classification

01. Fill in the blanks-

Metal oxide + Acid →

- Metal oxides are _____
_____ in Nature.

Acid + Metal → Salt +
_____ (g)

- $\text{Zn} + \text{HCl} \rightarrow$ _____
- $\text{Zn} + \text{HNO}_3 \rightarrow$ _____

Neutralisation Reaction

- $\text{NaOH} + \text{HCl} \rightarrow$ _____



Why do acids glow bulbs?

Is CH_3COOH or H_2SO_4
produce same glow?

Acid + _____
→ Salt + H_2O + CO_2

- $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow$ _____
- $\text{MgHCO}_3 + \text{HCl} \rightarrow$ _____

02. Answer the followings-



H_2 gas Test -











What happens if excess CO_2 is
passed?

What colour change occurs and
what new substances are formed?

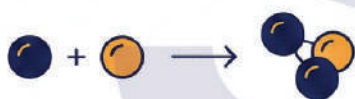
Lime water formula

D. Chemical Reactions

01. Write **P** for Physical Change & **C** for Chemical Change

 Melting of ice	 Freezing water	 Burning of iron	 Cooking of food
 Formation of precipitate	 Evolution of gas	 Heating copper sulphate	 Mixing solutions

02. What is a Chemical Reaction?



03. Indicators

Factor	Example
State	Burning of paper
Color	Iron in Copper Sulphate

04. Representation
